

## Solutions Chemical Thermodynamics

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Chemical thermodynamics - Wikipedia

Given that the free energy of formation of liquid water is  $-237 \text{ kJ / mol}$ , calculate the potential for the formation of hydrogen and oxygen from water. To

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solve this problem we must first calculate  $\Delta G$  for the reaction, which is  $-2(-237 \text{ kJ/mol}) = 474 \text{ kJ/mol}$ . Knowing that  $\Delta G = -nFE^\circ$  and  $n = 4$ , we calculate the potential is  $-1.23 \text{ V}$ .

Thermodynamics: Problems and Solutions | SparkNotes

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Problems And Solutions In Thermodynamics

The propensity for any two substances to form a solid solution is a complicated matter involving the chemical and crystallographic properties of the substances as well as temperature and pressure. At higher temperatures substances have greater atomic vibration and more open structures which are easier to distort locally to accommodate differently sized cations.

Solutions | Chemical Thermodynamics | Taylor & Francis Group

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