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2 POGIL™ Activities for High School Chemistry 3. In the space under each reaction in Model 1, write the oxidation number for every atom. Divide the work among your group members. An example is shown here:  $4\text{Fe}(s) + 3\text{O}_2(g) \rightarrow 2\text{Fe}_2\text{O}_3(s)$     0 0 +3 -2 0 +3 -2 -2 4. Identify any elements that changed

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oxidation number in the reactions in Model 1.

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LEO = Loss of Electrons is Oxidation 1 1. Consider the incomplete half-reactions below. a. Use oxidation numbers to identify the reactions below as oxidation or reduction. b. Place the correct number of electrons on

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the appropriate side of the reaction to complete the equation. PEOVCTI&..J 6 x STOP CF. e- 12. Consider Reaction A in Model 2.

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